

SVERIGE (A) ALLMÄNT TILLGÄNGLIG

(22) ANS DAT 1998-06-29 (21) ANS NR 9802303-9

(24) LÖPDAT 1998-06-29 ROTEL 333

(51) KLASSE B01J 20/18

B01J 20/32

(41) OFF DAT 1999-12-30 (74) OMBUD

(71) SÖKANDE Johan Sterte
 976 31 Luleå SE
 Lubomira Tosheva
 977 52 Luleå SE
 Valentin Valtchev
 976 31 Luleå SE
 Svetlana Mintova
 976 31 Luleå SE

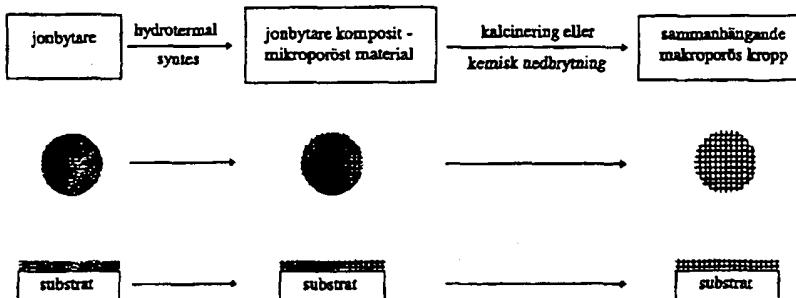
(72) UPPFINNARE Johan Sterte Luleå SE, Lubomira Tosheva Luleå SE, Valentin Valtchev Luleå SE, Svetlana Mintova Luleå SE

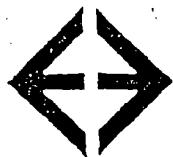
(30) PRIORITETSUPPGIFTER

(54) BENÄMNING Förfarande för framställning av makrostrukturer av mikroporösa material

(57) SAMMANDRAG

Uppfinningen avser ett förfarande för framställning av makrostrukturer av mikroporösa material. Förfarandet kännetecknas av att groddar bildade i eller genom jonbyte eller adsorption införda i en porös organisk jonbytare med önskad storlek, form och porositet, förmås växa och bilda en sammanhängande struktur genom vidare deposition av oorganiskt material från en synteslösning under hydrotermala betingelser samt att den organiska jonbytaren avlägsnas genom kemisk destruktion eller upplösning och därvid efterlämnar en oorganisk mikroporös struktur med storlek och form som den använda organiska jonbytaren.





RALPH McELROY TRANSLATION COMPANY

EXCELLENCE WITH A SENSE OF URGENCY®

TECHNICAL
TRANSLATIONS

1-800-531-9977
(512) 472-6753
FAX (512) 472-4591
FAX (512) 479-6703

ALL LANGUAGES

DATE:

4-16-99

TO:

Susan M. Mahnke
281-834-2911

FROM:

Laura Ann Lang

SENDING: COVER PAGE PLUS 32 PAGE(S)

MESSAGE: We are faxing you your translation. If you have any questions, please refer to our code number: 417-69097. The hard copy of this translation, along with the original, will be sent to you by first class mail.

If you prefer another method, please call this office immediately. Thank you for your business, and please do not hesitate to let us know if we may be of further service.

910 WEST AVE.
AUSTIN, TEXAS 78701
1-800-531-9977



(512) 472-6753
FAX (512) 472-4591
FAX (512) 479-6703

EXXON CHEMICAL COMPANY



BCTT Law Technology

VIA FAXSIMILE
512/472-4591

April 12, 1999

Ralph McElroy Translation Company
910 West Avenue
Austin, Texas 78701

Re: Translation of Swedish Patent Application 98082303-9
Billing Job No. 813,111

Dear Vicky:

Further to our phone conversation this morning, please find enclosed a copy of the referenced Swedish patent application that we would like to have translated into English by Friday, April 16, 1999 at the latest. Please let me know if you need a hard copy in the mail today and if you will be able to meet the requested date. ✓ Valfr VM m

Thank you very much for your assistance in this matter.

Very truly yours,

A handwritten signature in cursive ink that appears to read "Susan Mohnke".

Susan G. Mohnke 41829
Staff Legal Assistant

Enclosure

RECD APR 12 1999

417-69097

4500 Bayway Drive, Baytown, Texas 77520-9728
P.O. Box 2149, Baytown, TX 77522-2149
Tel: (281) 834-2136 Fax: (281) 834-2911

A Division of Exxon Corporation



[Patent entitled Process for Production of Macrostructures of a
Microporous Material]

Code: 417-69097

Ref. No.: 813.111

Title:

PROCESS FOR PRODUCTION OF MACROSTRUCTURES OF A MICROPOROUS MATERIAL

Technical area

The present invention concerns a process for production of macrostructures with controlled size and shape of a microporous inorganic material.

Prior art

Microporous material is characterized by a large specific surface area in pores with a pore radius below 20 Å and is used in a large number of applications of considerable commercial importance. In most of these applications, the fact that the phase interface between the solid porous material and the medium (liquid or gas) in which it is used is large is of decisive importance. Heterogeneous phase catalysts for refinery processes, for petrochemical conversion processes and for different environmentally related applications are often based on microporous material. Adsorbents for selective adsorption in the gas or liquid phase are microporous materials, like most of the inorganic ion exchangers used for selective separation of ionic compounds. In addition to these large-scale and relatively established applications, microporous materials have recently

become increasingly interesting in a number of more technologically advanced areas. Examples include use in chemical sensors, in fuel cells and batteries, in membranes for separation or catalytic purposes, during chromatography for preparative or analytical purposes, in electronics and optics and in the production of different types of composites.

Although a large phase interface is often a fundamental requirement for use of porous materials in different applications, a number of additional requirements related to the specific area of application are imposed on these materials. The large phase interface available in the micropores must be accessible and useable. Porosity, pore size and pore size distribution in large pores (meso- and macropores) are therefore often of major significance, especially when mass transport can be a rate-limiting factor. The chemical properties on the surface of the porous material can also be of decisive importance for the performance of the material in a given application. In this context, the purity of the material is, consequently, also significant. In most practical applications, the size and shape of the porous macrostructures and the degree of variation of these properties are of decisive importance. During use, the size and shape can influence properties like mass transport within the porous structures, pressure drop over a bed of particles of the material and the mechanical and thermal strength of the material. Which factor or factors are most important will vary strongly between different applications and are also highly dependent on the layout of the process in which the application occurs. Techniques that permit production of a material with increased specific surface area, pore structure (pore size/pore size distribution), chemical

composition, mechanical and thermal strength, as well as increased and uniform size and shape, are consequently required to tailor porous inorganic macrostructures to different applications. Microporous materials can be divided into crystalline, molecular sieve and amorphous materials. Molecular sieves are characterized by the fact that they have a pore system through their regular crystal structure, in which the pores have a very well defined size in the range 2-20 Å with an exact value determined by the structure. The size of most molecules that are gases and liquids at room temperature, both inorganic and organic, is found within this size range. By selecting a molecular sieve with the appropriate pore size, use of molecular sieves for separation of one substance (one type of molecule) in a mixture is made possible by selective adsorption, hence the name molecular sieve. In addition to selective adsorption of uncharged substances, the well defined micropore system of the molecular sieve offers a possibility for selective ion exchange of charged species and size-selective catalysis. In this case, properties other than the micropore structure in molecular sieves are also of major significance, like ion exchange capacity or specific surface area and acidity. Molecular sieves can be divided into a number of subgroups, depending on chemical composition and structure. A commercially important subgroup are the zeolites, which, by definition, are crystalline microporous aluminosilicates. Another interesting subgroup are the microporous metal silicates, which are structural analogs of the zeolites, but do not contain any (or very little) aluminum.

An excellent summary of the prior art, in terms of production, modification and characterization of molecular

sieves, is offered in the book Molecular Sieves - Principles of Synthesis and Identification (R. Szostak, Blackie Academic & Professional, London, 1998, Second Edition). In addition to molecular sieves, amorphous microporous materials, chiefly silica, aluminum silicate and aluminum oxide, have been used as adsorbents and catalyst supports. A number of long-known techniques, like spray drying, prilling, pelletizing and extrusion, have been and are being used to produce macrostructures in the form of, for example, spherical particles, extrudates, pellets and tablets of both micropores and other types of porous materials for use in catalysis, adsorption and ion exchange. A summary of these techniques is offered in Catalyst Manufacture, A. B. Stiles and T. A. Koch, Marcel Dekker, New York, 1995.

Because of limited possibilities with the known technique, considerable investment has been made to find new ways to produce macrostructures of microporous materials, with a certain emphasis on those in the form of films.

EP 94/01301 describes production of films of molecular sieves by a process in which seed crystals of molecular sieves are deposited on a substrate surface and then made to grow together into a continuous film. GB 94/00878 describes production of films of molecular sieves by introduction of a substrate to a synthesis solution adjusted for zeolite crystallization and crystallization with a gradual increase in synthesis temperature. SE 93/00715 describes production of colloidal suspensions of identical microparticles of molecular sieves with an average size below 200 nm. SE 90/00088 describes a method for production of an adsorbent material in the form of a monolith by impregnation of the monolithic cell structure with

a hydrophobic molecular sieve, followed by partial sintering of the molecular sieve with the material from which the cell structure is constructed.

Technical problem

Although a number of different techniques already exist for production of microporous inorganic macrostructures with the desired size and shape, these techniques are beset with a number of limitations that affect the properties of the macrostructures during use in the intended application. Most of these techniques require the use of a binder to give the structure acceptable mechanical strength. This binder often adversely affects other desired properties, like high specific surface area and uniform chemical composition. For most of the existing techniques, other possibilities for keeping variations in size and shape within narrow limits are sharply constrained. If a well defined size is desired with a narrow particle size distribution, one is most often obliged to carry out processing by separation, which leads to considerable waste during manufacture. The use of different types of binders also affects the pore structure in the resulting macrostructure and it is often necessary to find a compromise, in which the desired pore structure is weighed against the mechanical properties of the material. It is often desirable to have a bimodal pore size distribution in the macrostructures of macroporous materials, in which the micropores maintain a large specific phase interface, whereas the larger pores in the meso- or macropore range permit transport of molecules to the surface and, in this way, prevent limitations caused by slow diffusion. During production of

microporous macrostructures according to the known technique, a secondary system of pores within the meso- and/or macropore range can be produced by admixing a particulate inorganic material or by admixing organic material (for example, cellulose fibers), which are later eliminated by calcining. Both of these techniques, however, produce an adverse effect on the other properties of the resulting material.

Solutions

With the present invention, it has been possible to eliminate the drawbacks of the known methods and devise a process for production of macrostructures of microporous materials with controlled size, shape and porosity according to the subsequent Claims 1-9.

One objective of the present invention was to eliminate the drawbacks in the known methods for production of microporous macrostructures and devise a new process that permits production of these structures without addition of binders and with a uniform final composition. Another objective of the present invention was to devise a process, according to which the final shape, size and size distribution of the structure can be controlled with high precision by selecting the starting material. Another purpose of the present invention was to devise a process according to which both the micropore structure of the material and a secondary system of larger pores can be controlled by selecting the starting material and synthesis conditions. It was also an objective of the present invention to devise a process for production of macrostructures of microporous material with good mechanical and thermal stability.

Figure description

Figure 1 shows a schematic description of the different stages in production of spherical particles or thin films of microporous material according to the invention.

Figure 2 shows adsorption-desorption isotherms measured for spherical particles of amorphous silica produced according to Application Examples 1 (I) and 2 (II) according to the invention.

Figure 3 shows a photograph taken with a scanning electron microscope, at two different magnifications, of spherical particles of a molecular sieve silicalite-1 produced according to the invention.

Figure 4 shows an X-ray diffraction pattern for spherical particles of the molecular sieve silicalite-1 produced according to the invention.

Detailed explanation of the invention

The present invention concerns a process for production of macrostructures of microporous material, characterized by the fact that seeds formed in or introduced by ion exchange or adsorption to a porous organic ion exchanger with the desired size, shape and porosity, are made to grow and form a continuous structure by further deposition of inorganic material from a synthesis solution under hydrothermal conditions, and that the organic ion exchanger is then eliminated by chemical destruction or dissolution and, in so doing, leaves behind an inorganic microporous structure with the size and shape of the employed

organic ion exchanger. A schematic description of the method according to the invention is shown in Figure 1.

Microporous material refers to a material in which most of the specific surface area is found in pores with a diameter below 20 Å. A crystalline microporous material, usually called molecular sieves or zeolites, is of special interest for the invention. Materials like silicalite-1, ZSM-5, zeolites A, Y, mordenite, Beta, L, ZSM-2, X and hydroxysodalite are of special interest. Different types of amorphous silicates are also of interest. Macrostructures refer to structures with a size that exceeds 0.01 mm in one dimension, preferably 0.1 mm and, in most cases, 1.0 mm. Examples of macrostructures are spherical particles, cylindrical extrudates, pellets, fibers, thin films applied to different forms of substrates and other composites, in which the microporous material is combined with other types of material. Organic ionic exchanger refers to a polymer-based porous material with a fixed surface charge and with high ion exchange capacity for anions or cations. A large number of different types of organic ion exchangers are now commercially available under the tradenames Dowex and Amberlite. This covers a broad spectrum of different cation and anion exchangers with varying ion exchange capacity, porosity, pore size and particle size. Ion exchangers with an apparent anion exchange capacity, typically >1 mEq per g of dry anion exchanger, are of special interest to the present invention. Ion exchangers of the strongly basic type, containing quaternary ammonium groups, have been found to be particularly suited for use according to the present invention. Commercially available ion exchangers are generally in the form of spherical particles with a relatively narrow particle size distribution. Organic ion exchange

materials with a size and shape other than spherical, for example, fibers or flakes, however, can be produced according to known techniques. It is also known that films of organic ion exchangers can be deposited on different forms of substrates.

Seeds refer to silicate species, metal silicate species or aluminosilicate species present in the synthesis mixtures or formed in an early stage of synthesis, characterized by the fact that by treatment in a solution with appropriate concentration and under suitable conditions, they can be made to grow and form a continuous structure in the pore system of the ion exchanger. The term "seeds" also includes microcrystals, i.e., crystals of molecular sieves with a size below 200 nm, and whose crystal structure can be identified by X-ray diffraction. Production of microcrystals of molecular sieves suitable as seeds for production of microporous macrostructures according to the invention is described in WO 94/05597.

In a second phase in production of microporous macrostructures according to the invention, the seeds formed on or bonded to the surface in the organic ion exchanger are made to grow by hydrothermal treatment in an appropriate synthesis solution. Through this growth a continuous three-dimensional network of microporous material is formed in the pore structure of the employed ion exchange structure. After this stage, the product is therefore a composite material, consisting of two continuous three-dimensional networks, one organic, consisting of the polymer structure of the ion exchanger, and one inorganic, consisting of the introduced inorganic microporous material. Introduction of seeds can be carried out physically in a separate stage, with a subsequent growth stage under hydrothermal conditions in a synthesis solution. However, it is

also possible and often advantageous not to separate these stages, but instead to directly introduce the ion exchanger material into a synthesis solution and expose this to hydrothermal conditions, during which seeds are formed in or ion-exchanged from the synthesis solution to the ion exchanger, to then grow into a continuous structure.

Molecular sieves of the zeolite or crystalline microporous metal silicate type are generally produced by hydrothermal treatment of a silicate solution with well defined composition. This composition and the synthesis parameters, like temperature, time and pressure, dictate which type of product is obtained and the size and shape in the crystals that are formed. This applies both in syntheses, in which the final product is deposited as crystals in the porous structure of an ion exchanger, and in conventional synthesis, when the final crystal size is most often much larger. The type of material deposited in the pore system of the ion exchanger is therefore dependent on the composition of the synthesis mixture and the conditions during synthesis. During crystallization of macrostructures of a given molecular sieve according to the present invention, it has been found suitable to use synthesis mixtures, which, in the absence of ion exchanger material, result in colloidal suspensions of the molecular sieve in question. In some cases, the ion exchanger material has also been found to influence the result of synthesis. For example, when certain ion exchanger materials are used, an amorphous macrostructure of microporous silica with very large specific surface area can be produced by means of a synthesis solution and under conditions which, in the absence of an ion exchanger, produce a crystalline molecular sieve. A parameter of great interest in this connection has been found to

be the ratio between amount of ion exchanger and amount of synthesis solution. Under otherwise identical conditions and during the use of the same synthesis solution, a low ratio between the amount of ion exchanger and amount of synthesis solution yields a well crystallized product, whereas an intermediate ratio yields a mixture of crystalline and amorphous microcrystalline material, and a high ratio leads to an entirely amorphous microporous material. Hydrothermal treatment refers to treatment in aqueous solution or aqueous suspension at a temperature exceeding 50°C, preferably exceeding 80°C and, in most cases, exceeding 95°C. The composite of ion exchanger and microporous inorganic material obtained after this process can be of interest by itself in certain commercial applications. However, for most potential areas of application it is advantageous to eliminate the organic ion exchanger structure. This can occur in a last stage of the process according to the invention, which leaves behind only a microporous material and a secondary pore system with a porosity and pore size caused by the structure of the employed organic ion exchanger. Elimination of the organic material that makes up the ion exchanger preferably occurs by calcining at a temperature exceeding 400°C in the presence of acid, in which this material is burned to mostly carbon dioxide and water. As an alternative, the organic material can be eliminated by selective dissolution with a solvent that dissolves the ion exchanger, but not the inorganic material, or with selective decomposition of the inorganic material by means of a chemical reaction other than a combustion reaction.

The resulting material according to the invention is a replica in size and shape of the organic ion exchanger used as

starting material. This means that the possibilities for controlling the size, shape and meso/macroporosity in the inorganic microporous material are largely determined by the possibilities of structural manipulation of the properties of the ion exchanger. The secondary pore structure that develops during removal of the organic ion exchanger material, however, can be further manipulated by continued deposition after removal of this material. By growth of the inorganic structure after this stage, the secondary pore structure can be more or less sealed and, in the extreme case, leave behind a homogeneous microporous material (without porosity in the meso/macropore range). This could be of interest, for example, in the production of thin films of microporous structures, for use in applications, like membranes for catalyst or separation purposes, or in chemical sensors. It is also possible, according to a known technique, to coat the surface of the macrostructures of a given type of microporous material produced according to the invention with a thin film of another type of material, something that could be of interest in a catalytic context or during use of macrostructures for controlled dosage of drugs or pesticides.

Characterization

The process according to the present invention was evaluated by means of investigation of materials produced according to the process with a scanning electron microscope (SEM), X-ray diffractometry (XRD), spectroscopy and by measurements of the specific surface area and pore size distribution with krypton or nitrogen adsorption.

Scanning electron microscope studies were conducted on samples coated with gold (by a sputtering technique). A scanning electron microscope of the Philips XL 30 type with an LaB₆ [sic] emission source was used in these studies.

X-ray diffraction studies were conducted with a Siemens D-5000 powder diffractometer.

Nitrogen adsorption measurements to determine specific surface area and particle size distribution were carried out with an ASAP 2010 from Micrometrics Instruments, Inc.

Elemental analysis concerning carbon, nitrogen and hydrogen was carried out on certain samples by means of an analytical instrument from LECO Corporation (LECO CHN-600). The particle size and particle size distribution for the colloidal suspensions of discrete microcrystals of molecular sieves used as starting material according to the process were investigated by dynamic light scattering (ZetaPlus, Brookhaven Instruments).

Application examples

The invention is explained below with some application examples. However, these are not to be viewed as limiting for the invention.

Example 1

This example describes production of spherical particles of microporous amorphous silica with very high specific surface area.

A synthesis solution with the composition (on a molar basis): 9 TPAOH/25 SiO₂/480 H₂O/100 EtOH (TPAOH =

tetrapropylammonium hydroxide, EtOH = ethanol) was produced by mixing 20.0 g tetraethoxysilane (>98%, Merck), 34.56 g tetrapropylammonium hydroxide (1.0M solution, Sigma) and 5.65 g distilled water. The mixture was allowed to hydrolyze in a polyethylene flask on a shaking table for 12 h at room temperature. 1.0 g of a strongly basic anion exchanger of the Dowex 1X2-100 type was added to 10 g of this synthesis solution. The anion exchanger was present as spherical particles in the particle size range 50-100 mesh (dry) and had an ion exchange capacity specified by the manufacturer of 3.5 mEq/g. The mixture of ion exchanger and synthesis solution was treated in a polyethylene reactor equipped with a reflux condenser in an oil bath at 100°C for 48 h. After this time, the ion exchanger particles were separated from the solution by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon they were separated again by means of filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 60°C for 12 h. The particles were finally calcined at 600°C in air for 4 h, after heating to this temperature at a rate of 10°C/min. The resulting material consisted of very hard, solid, white spherical particles with a size distribution identical to that in the employed ion exchanger. Elemental analysis showed that the particles were almost entirely free of carbon, hydrogen and nitrogen, which clearly shows that the ion exchanger material had been completely eliminated in the calcining stage. X-ray diffractometry also showed that the material was completely amorphous. The particles were further analyzed by nitrogen adsorption measurements at the boiling point of nitrogen to

determine the specific surface area, the adsorption isotherm and pore size distribution. The specific surface area was calculated from the adsorption data according to the BET equation as 1220 m²/g. The recorded isotherm was of type I, see Figure 2, which is typical of microporous materials. Calculation of the pore size distribution by the BJH method (desorption isotherm) showed that a very small fraction (about 20 m²/g) of the total specific surface area of the material was found in pores in the mesopore range (diameter >20 Å). The average pore diameter was calculated at 9.5 Å by the Horvath-Kawazoes method.

Example 2

This example describes production of spherical particles of amorphous aluminum silicate with high specific surface area in pores in both the micro- and mesopore range.

25 g of a synthesis solution with the molar composition: 2.4 Na₂O/1.0 TEACl/0.4 Al₂O₃/10 · SiO₂/460 H₂O (TEACl = tetraethylammonium chloride) were added to 2.0 g of a strongly basic ion exchanger of the type Dowex MSA-1 (particle size 20-50 mesh (dry); ion exchange capacity: 4 mEq/g) in a polyethylene reactor. The synthesis mixture was prepared by first dissolving 0.75 g sodium aluminate (50.6 wt% Al₂O₃, 36 wt% Na₂O) in 35 g of a 1M NaOH solution at 100°C. This solution was then added to a mixture of 40 g distilled water, 1.66 g TEACl (Merck) and 15 g silica sol (Bindzil 40/130, Eka Chemicals AB, solids content 41.36 wt%, 0.256 wt% Na₂O) during agitation for 2 h. The mixture of ion exchanger and synthesis solution was treated in a polyethylene reactor equipped with a reflux condenser in an oil bath at 100°C for 48 h. After this time, the

ion exchanger particles were separated from the solution by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon they were separated again by filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 60°C for 12 h. The particles were finally calcined at 600°C in air for 4 h, after heating to this temperature at a rate of 10°C/min.

Visual inspection and analysis with a scanning electron microscope showed that the resulting material consisted of very hard, solid, white spherical particles with size distribution identical to that in the employed ion exchanger. Elemental analysis showed that the particles were almost entirely free of carbon, hydrogen and nitrogen, which clearly shows that the ion exchanger material had been fully eliminated in the calcining stage. X-ray diffractometry also showed that the material was completely amorphous. The particles were further analyzed by nitrogen adsorption measurements at the boiling point of nitrogen to determine the specific surface area, adsorption isotherms and pore size distribution. The specific surface area was calculated from the adsorption data according to the BET equation as 594 m²/g. The recorded isotherm was of type IV, see Figure 2. Calculation of the pore size distribution by the BJH method (desorption isotherm) showed that a relatively large percentage of the total (cumulative) pore volume (about 65%) was found in pores in the mesopore range (radius >20 Å).

Example 3:

This example describes production of spherical particles of molecular sieve silicalite-1 with the process according to the invention.

14.3 g of a synthesis solution with the molar composition: 9 TPAOH/25 SiO₂/480 H₂O/100 EtOH were added to 1.0 g of macroporous strongly basic ion exchanger of the Dowex MSA-1 type (particle size 20-50 mesh (dry); ion exchange capacity: 4 mEq/g). The synthesis mixture was prepared as described in Example 1. The mixture of ion exchanger and synthesis solution was treated in a polyethylene reactor equipped with a reflux condenser in an oil bath at 100°C for 48 h. After this time, the ion exchanger particles were separated from the solution and the material was crystallized in the bulk phase by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon they were separated again by filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 60°C for 12 h. The particles were finally calcined at 600°C in air for 10 h, after heating to this temperature at a rate of 1°C/min.

Visual inspection and scanning electron microscopy revealed that the resulting material consisted of very hard, solid (homogeneous), white spherical particles with a size distribution identical to that in the employed ion exchanger. Elemental analysis showed that the particles were almost entirely free of carbon, hydrogen and nitrogen, which clearly shows that the ion exchanger material was fully eliminated in the calcining stage. Figure 3 shows an SEM photograph taken at

two different magnifications of a product according to this example. The photo taken at the lower magnification shows the spherical character of the particles, whereas the one taken at high magnification shows the presence of small primary particles (primary crystals) with a size of about 100 nm. X-ray diffractometry revealed that the material is crystalline and consists of silicalite-1, but that it also contains a percentage of amorphous material. An X-ray diffraction pattern for this sample is shown in Figure 4. Analysis with nitrogen adsorption gave a specific surface area of 438 m²/g and showed that most of the pore volume was found in micropores with an average pore diameter of 6 Å, calculated according to the Horvath-Kawazoe method.

Example 4

This example describes production of spherical particles of molecular sieve zeolite ZSM-5 with the process according to the invention.

15 g of a synthesis solution with the molar composition: 0.58 Na₂O/9 TPAOH/0.5 Al₂O₃/25 SiO₂/405 H₂O were added to 1.0 g of a macroporous strongly basic anion exchanger of the Dowex MSA-1 type (particle size 20-50 mesh (dry); ion exchange capacity: 4 mEq/g). The synthesis mixture was prepared by first dissolving 0.408 g aluminum isopropoxide (Sigma) in 10 g of 1.0M tetrapropylammonium hydroxide (Sigma). Another solution was prepared by dissolving 6.0 g freeze-dried silica sol (Bindzil 30/220, 31 wt% SiO₂, 0.5 wt% Na₂O Eka Chemicals, AB) in 26 g 1.0M TPAOH at 100°C. The two solutions were mixed during agitation for 30 min. The mixture of ion exchanger and synthesis solution was

treated in a polyethylene reactor equipped with a reflux condenser in an oil bath at 100°C for 20 days. After this time, the ion exchanger particles were separated from the solution and the material was crystallized in the bulk phase by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon they were separated again by filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 50°C for 12 h. The particles were finally calcined at 600°C in air for 10 h, after heating to this temperature at a rate of 1°C/min.

Visual inspection and analysis with a scanning electron microscope showed that the product largely consisted of white, solid particles with a size and shape identical to that of the employed ion exchanger. A relatively large fraction of the product, however, was shown to consist of particles with roughly the same size as the employed ion exchanger, but with a more irregular shape. SEM analysis at high magnification showed that the particles consist of intergrown crystals with a morphology typical of MFI structures and with a size of about 1 μm . X-ray diffractometry showed that the particles consist of zeolite ZSM-5, but also a relatively large fraction of amorphous material. The specific surface area was measured by nitrogen adsorption at 612 m^2/g .

Example 5

This example describes production of spherical particles of molecular sieve zeolite A with the process according to the invention.

18.0 g of a synthesis solution with the molar composition: 0.22 Na₂O/5.0 SiO₂/Al₂O₃/8 TMA₂O/400 H₂O were added to 1.0 g of a strongly basic anion exchanger of the Dowex MSA-1 type. The synthesis mixture was prepared by first dissolving 1.25 g aluminum isopropoxide (Sigma) and 9.0 g tetramethylammonium hydroxide pentahydrate (Sigma) in 0.90 g of 1.0M solution of NaOH and 3.0 g water during agitation for 2 h. This solution was added, in turn, to a mixture of 3.0 g silica sol (Bindzil 30/220, see Example 4) and 12 g of distilled water, whereupon the resulting solution was agitated for 3 h. The mixture of ion exchanger and synthesis solution was treated in a polyethylene reactor equipped with a reflux condenser in an oil bath at 100°C for 10 h. After this time, the ion exchanger particles were separated from the solution and the material was crystallized in the bulk phase by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon they were separated again by filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 60°C for 12 h. The particles were finally calcined at 600°C in air for 10 h, after heating to this temperature at a rate of 1°C/min.

Visual inspection and analysis by scanning electron microscopy showed that the product largely consisted of light brown, solid particles with a size and shape identical to that of the employed ion exchanger. A smaller fraction of the product consisted of fragmented particles. SEM at high magnification showed that the particles are homogeneous and constructed from intergrown primary particles with a size up to about 300 nm. X-ray diffractometry showed that the resulting material

contained zeolite A and a certain amount of amorphous material. Nitrogen adsorption measurements gave a specific surface area (according to the BET equation) of 306 m²/g and indicated the presence of both micro- and mesoporosity.

Example 6

This example describes production of spherical particles of molecular sieve zeolite Beta with the process according to the invention.

15 g of a synthesis solution with the molar composition: 0.35 Na₂O/9 TEAOH/0.5 Al₂O₃/25 SiO₂/295 H₂O were added to 1.0 g of a strongly basic anion exchanger of the Dowex MSA-1 type. The synthesis mixture was prepared by dissolving 0.81 g aluminum isopropoxide (Sigma) in 6.0 g tetraethylammonium hydroxide (TEAOH, 20% solution) at 100°C. This solution was added to a solution of 6.0 g freeze-dried silica sol (Bindzil 30/220, see Example 4) dissolved in 20 g TEAOH (20% solution) and the resulting solution was agitated for 30 min. The mixture of ion exchanger and synthesis solution was treated in a polyethylene reactor equipped with a reflux condenser in an oil bath at 100°C for 8 days. After this time, the ion exchanger particles were separated from the solution and the material was crystallized in the bulk phase by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon the particles were separated again by filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 60°C for 12 h. The particles were finally

calcined at 600°C in air for 10 h, after heating to this temperature at a rate of 1°C/min.

Visual inspection, as well as analysis with a scanning electron microscope, showed that the product largely consisted of hard, white, solid particles with a size and shape identical to that of the employed ion exchanger. SEM analysis at high magnification shows that the material is constructed of intergrown primary particles with a size of about 80 nm. X-ray diffractometry showed that the particles contained zeolite Beta as the only crystalline phase. The specific surface area calculated with the BET equation, based on nitrogen adsorption data, was 580 m²/g.

Example 7

This example describes how a film of molecular sieve silicalite-1 can be built up on the surface of a macrostructure of silicalite-1 produced according to Example 3.

10.0 g of synthesis solution with the composition and preparation according to Example 3 were added to 0.20 g of calcined product produced according to Example 3. This mixture was treated at 100°C in a polyethylene reactor equipped with a reflux condenser for 48 h. After this time, the particles were separated from the solution and the material was crystallized in the bulk phase by filtration and treated in a 0.1M ammonia solution in an ultrasound bath for 15 min, whereupon they were separated again by filtration. The particles were finally washed three times by suspension in distilled water, followed by separation by filtration, and then dried in a heating cabinet at 60°C for 12 h. Part of the material was calcined at 600°C for

10 h, after heating to this temperature at a rate of 1°C/min. X-ray diffraction measurements on the calcined sample revealed that the sample contained silicalite-1 as the only crystalline phase. Scanning electron microscopy detected an outer layer of silicalite-1 on the surface of the particles, a layer that synthesis had built up from about 300-nm large primary particles. The specific surface area was determined for the uncalcined sample as 92 m²/g, whereas the corresponding value measured for the calcined sample was 543 m²/g. The difference in the surface before and after calcining indicates that the outer shell of silicalite-1 effectively encloses the open pore system in the original particles.

Claims

1. Process for production of macrostructures of microporous material, characterized by the fact that seeds formed in or introduced by ion exchange or adsorption to a porous organic ion exchanger with the desired size, shape and porosity are made to grow and form a continuous structure by further deposition of inorganic material from a synthesis solution under hydrothermal conditions, and that the organic ion exchanger is eliminated by chemical destruction or dissolution and, in so doing, leaves behind an inorganic microporous structure with the size and shape of the employed organic ion exchanger.

2. Process according to Claim 1, characterized by the fact that the organic ion exchanger is an anion exchanger.

3. Process according to any of Claims 1-2, characterized by the fact that the seeds are oligomeric anions of silicates or microcrystals of a molecular sieve.

4. Process according to any of Claims 1-3, characterized by the fact that the synthesis solution has a composition, which, during hydrothermal treatment in the absence of ion exchanger, results in crystallization of a molecular sieve.

5. Process according to any of Claims 1-4, characterized by the fact that the organic ion exchanger structure is entirely or partially eliminated by combustion.

6. Process according to any of Claims 1-5, characterized by the fact that the microporous material is amorphous silica or amorphous aluminosilicate with a specific surface area exceeding 200 m²/g.

7. Process according to any of Claims 1-5, characterized by the fact that the microporous material is a crystalline molecular sieve.

8. Process according to any of Claims 1-5 and 7, characterized by the fact that the molecular sieve is represented by silicalite, hydroxysodalite, Ti-silicalite, mordenite or any of the zeolites: A, Beta, L, X, Y, ZSM-5 or ZSM-2.

9. Process according to any of Claims 1-8, characterized by the fact that the synthesis solution is a silicate, aluminosilicate or metal silicate solution.

Abstract

The invention concerns a process for production of macrostructures of a microporous material. The process is characterized by the fact that seeds formed in or introduced by ion exchange or adsorption to a porous organic ion exchanger with the desired size, shape and porosity are made to grow and

form a continuous structure by further deposition of inorganic material from a synthesis solution under hydrothermal conditions, and that the organic ion exchanger is eliminated by chemical destruction or dissolution and, in so doing, leaves behind an inorganic microporous structure with the size and shape of the employed organic ion exchanger.

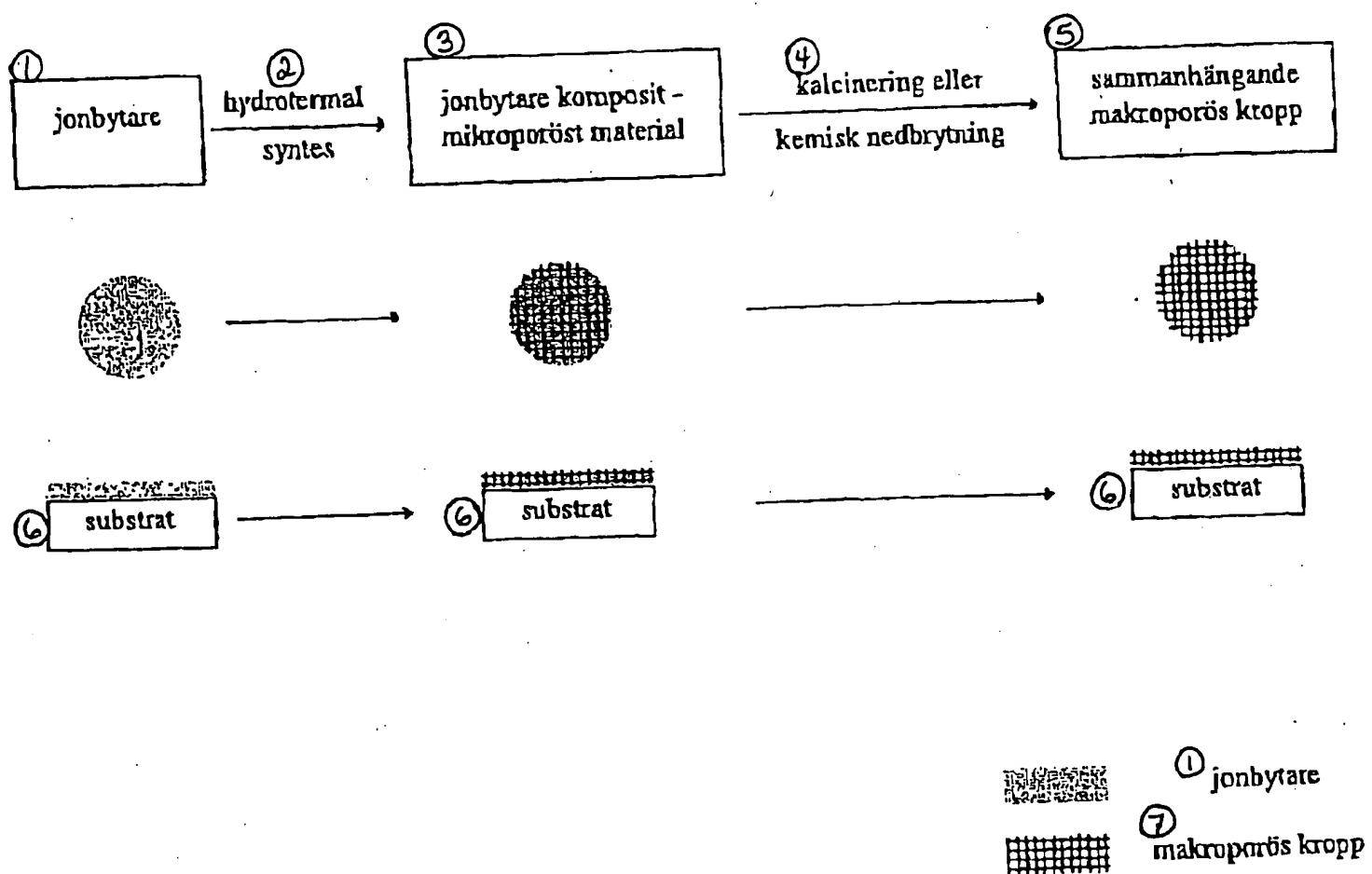


Figure 1

Key: 1 Ion exchanger
2 Hydrothermal synthesis
3 Ion exchanger composite-microporous material
4 Calcining or chemical decomposition
5 Continuous macroporous body
6 Substrate
7 Macroporous body

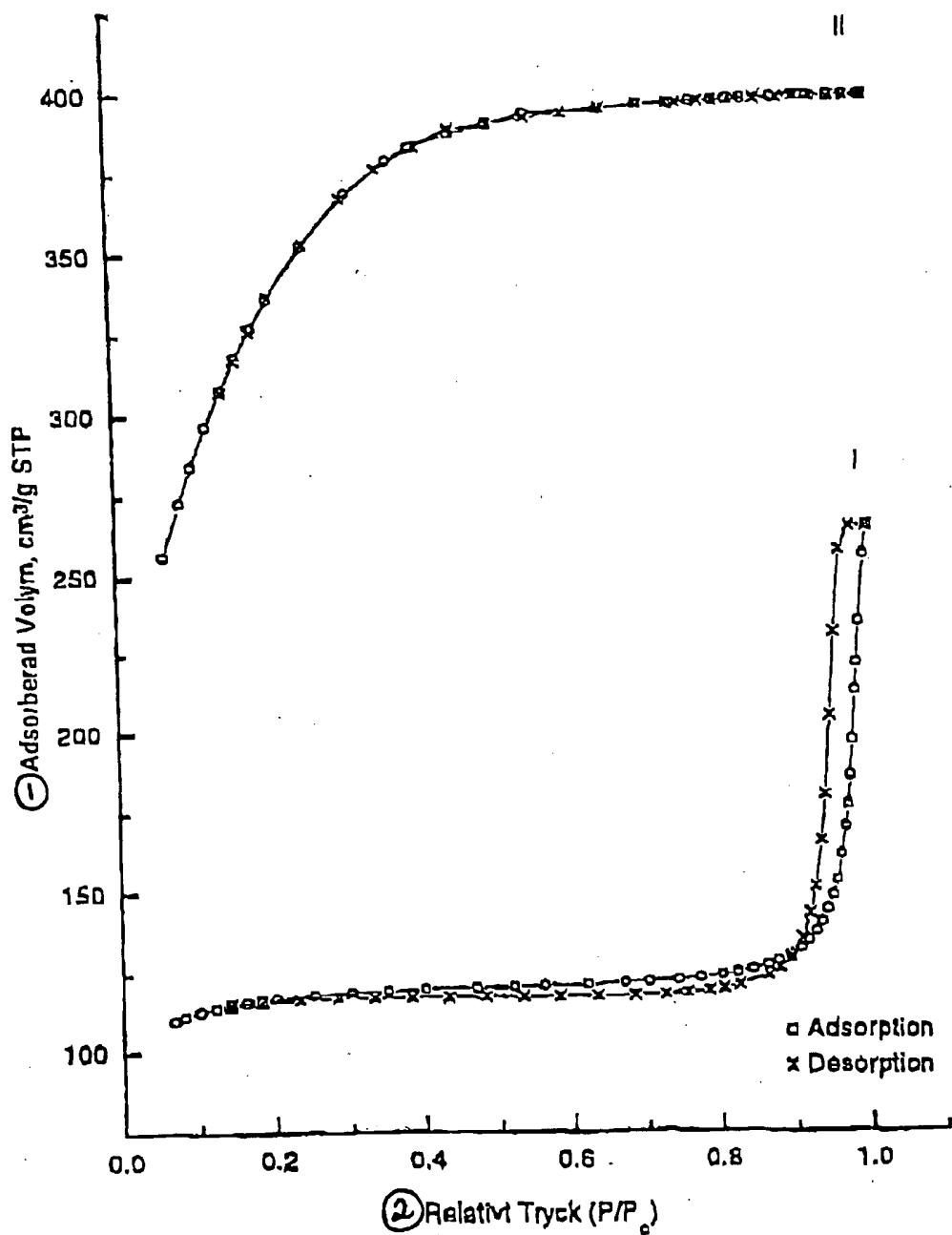


Figure 2

29

Key: 1 Adsorbed volume, cm^3/g STP
2 Relative pressure (P/P_0)

30

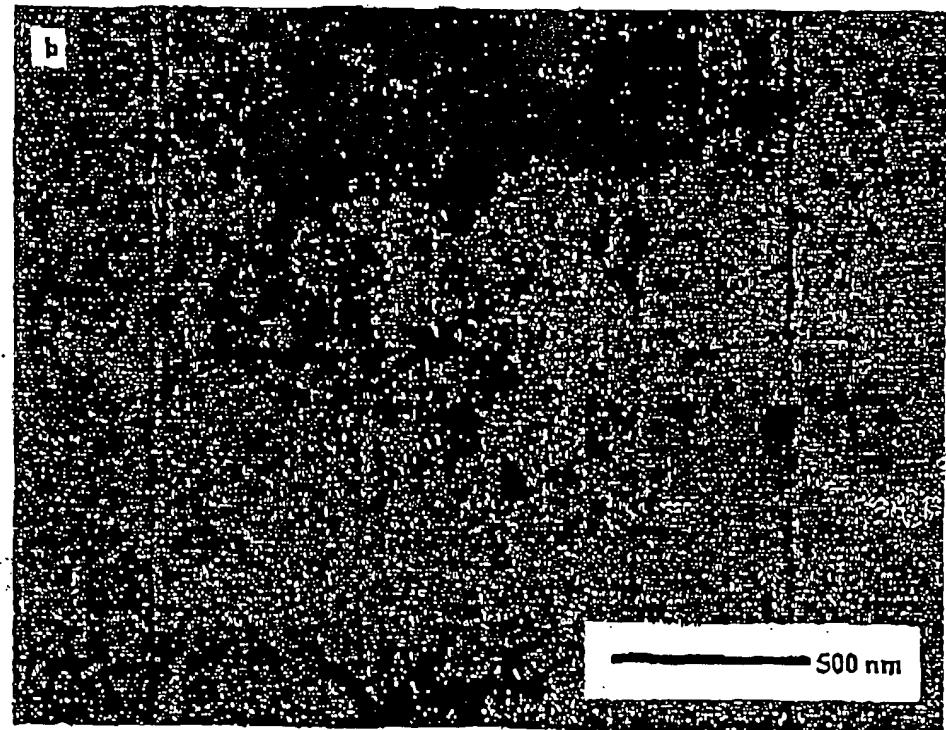
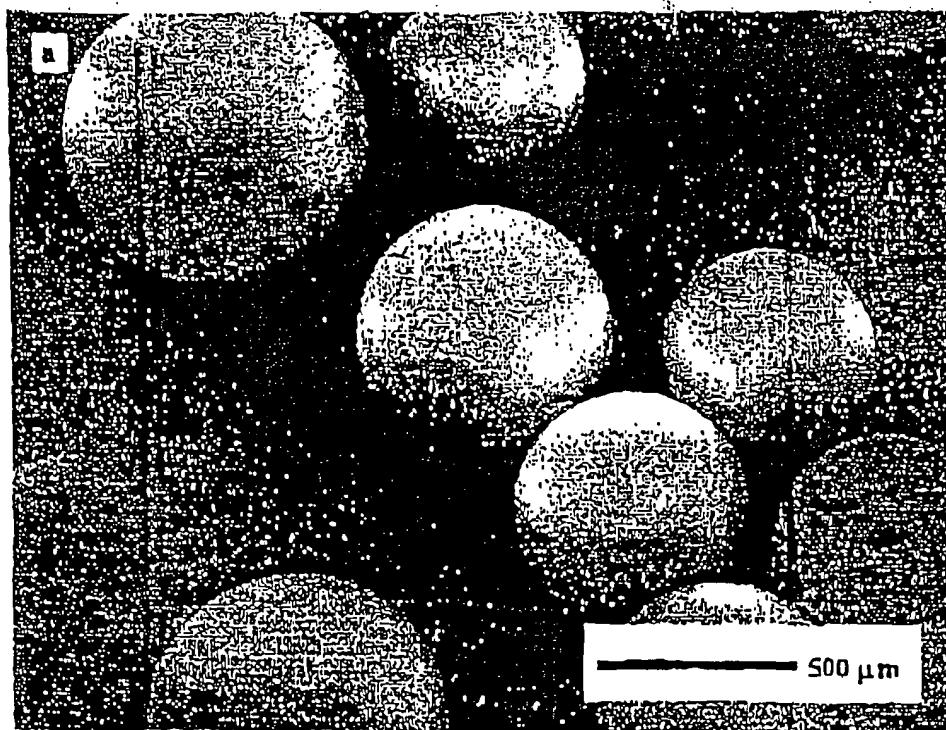


Figure 3

31

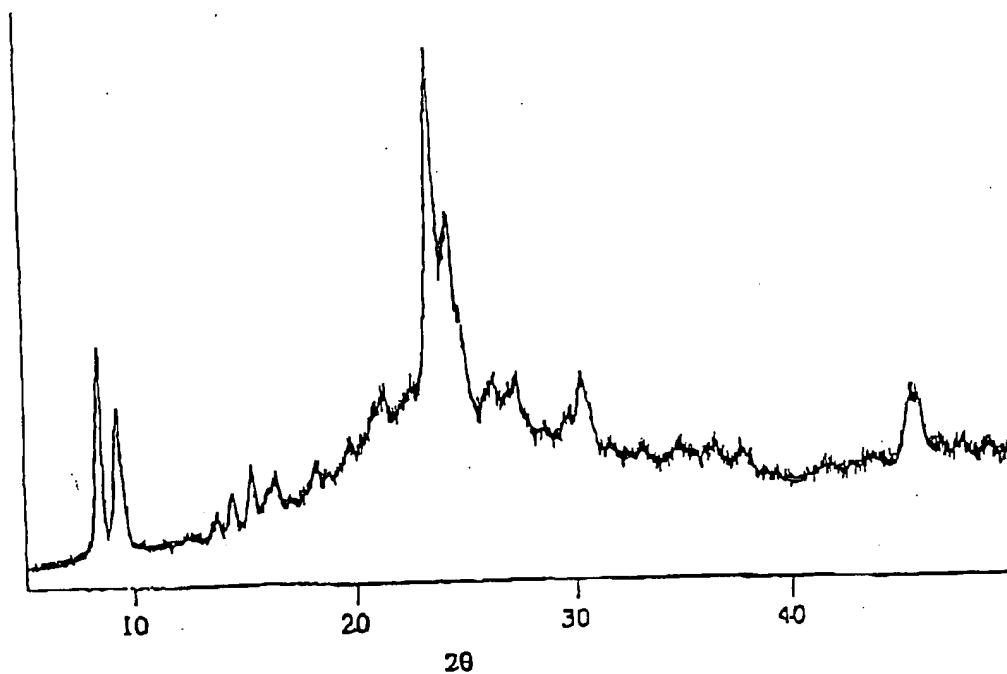


Figure 4